

REVISION PRACTICE ASSIGNMENT (RPA)

SUBJECT- CHEMISTRY

SESSION-2020-21

CLASS-XII

TOPIC: SOLUTION

Q. Answer the following question: -

1. Which of the following units is useful in relating concentration of solution with its vapour pressure.
a) mole fraction b) ppm c) mass percentage d) molality
2. Number of moles of solute present in one litre of solution is known as _____.
3. A solution is said to be _____ if it obeys Raoult's law at all concentration and temperature.
4. Those properties of solution that depend upon number of solute particles are known as _____.
5. Binary solutions which have the same composition in liquid phase as well as vapour phase and boil at constant temperature are known as _____.

Q. Very short answer type question: -

- 1) Write the unit of molarity?
- 2) What will happen to the boiling point of solution on mixing two miscible liquids showing negative deviation from Raoult's law?
- 3) Can a solution of its own have osmotic pressure?
- 4) What is common in all the four colligative properties?
- 5) What type of azeotropes will result on mixing chloroform and acetone?

Q. Short answer type question: -

- 1) Why it is advisable to use ethylene glycol in car radiator during summer?
- 2) Differentiate between molarity and molality of a solution.
- 3) Draw the graph that shows positive deviation from Raoult's law.

Q. - Long answer type question: -

- 1) a) Define molality and mole fraction
b) Concentrated nitric acid used in laboratory work is 68% nitric acid by mass in aqueous solution. What should be the molarity of such sample of acid if density of solution is 1.504 gram per ml?